

Chapter 4, Lesson 2: The Periodic Table

Key Concepts

- The periodic table is a chart containing information about the atoms that make up all matter.
- An element is a substance made up of only one type of atom.
- The atomic number of an atom is equal to the number of protons in its nucleus.
- The number of electrons surrounding the nucleus of an atom is equal to the number of protons in its nucleus.
- Different atoms of the same element can have a different number of neutrons.
- Atoms of the same element with different numbers of neutrons are called “isotopes” of that element.
- The atomic mass of an element is the average mass of the different isotopes of the element.
- The atoms in the periodic table are arranged to show characteristics and relationships between atoms and groups of atoms.

Summary

Students will begin to look closely at the periodic table. They will be introduced to the basic information given for the elements in most periodic tables: the name, symbol, atomic number, and atomic mass for each element. Students will focus on the first 20 elements. They will try to correctly match cards with information about an element to each of the first 20 elements. Students will then watch several videos of some interesting chemical reactions involving some of these elements.

Objective

Students will identify different atoms by the number of protons in the nucleus and realize that the number of electrons equals the number of protons in a neutral atom. They will also be able to explain the meaning of atomic number and atomic mass.

Evaluation

The activity sheet will serve as the “Evaluate” component of each 5-E lesson plan. The activity sheets are formative assessments of student progress and understanding.

About this Lesson

Lessons 2 and 3 both use the 20 atom description cards beginning on page 240.

Teacher preparation

Print out the 20 pages of element cards. The first page is shown. Laminate each page and cut out the cards. For Lesson 2, you will need the 5 cards for each element from the left side of each sheet. You will also need the card in the upper right corner. This is the atom name card. Tape each of the 20 atom name cards to a spot in the room where students can place the cards that match that atom nearby. For Lesson 3, you will need the atom name card, taped in the same location in the room, and the four cards beneath it. Divide the class into 10 groups of 2 or 3 students each.

The atom you are looking for has: 1 Proton in its Nucleus.	Atomic Number 1 HYDROGEN (H) Atomic Mass: 1.01
The atom you are looking for has: 1 Electron surrounding its Nucleus.	The atom you are looking for has the same energy level!
The atom you are looking for has: 0 Neutrons (usually) in its Nucleus.	The atom you are looking for has Neutrons on the First energy level!
The atom you are looking for has: 1 fewer Proton than Helium (He).	The atom you are looking for is directly above the atom with the same level.
The atom you are looking for has: 2 fewer Electrons than Lithium (Li).	The atom you are looking for is the only atom with only 1 Electron on the First energy level!

ENGAGE

1. Introduce students to the periodic table.

Project the image *Periodic Table*.

www.middleschoolchemistry.com/multimedia/chapter4/lesson2#periodic_table

Tell students that this is the periodic table. Explain that each box contains information about a different atom. The periodic table shows all the atoms that everything in the known universe is made from. It's kind of like the alphabet in which only 26 letters, in different combinations, make up many thousands of words. The 100 or so atoms of the periodic table, in different combinations, make up millions of different substances.

The Periodic Table of the Elements

1 H Hydrogen 1.01																	2 He Helium 4.00				
3 Li Lithium 6.94	4 Be Beryllium 9.01															5 B Boron 10.81	6 C Carbon 12.01	7 N Nitrogen 14.01	8 O Oxygen 16.00	9 F Fluorine 19.00	10 Ne Neon 20.18
11 Na Sodium 22.99	12 Mg Magnesium 24.31															13 Al Aluminum 26.98	14 Si Silicon 28.09	15 P Phosphorus 30.97	16 S Sulfur 32.07	17 Cl Chlorine 35.45	18 Ar Argon 39.95
19 K Potassium 39.10	20 Ca Calcium 40.08	21 Sc Scandium 44.96	22 Ti Titanium 47.87	23 V Vanadium 50.94	24 Cr Chromium 52.00	25 Mn Manganese 54.94	26 Fe Iron 55.85	27 Co Cobalt 58.93	28 Ni Nickel 58.69	29 Cu Copper 63.55	30 Zn Zinc 65.39	31 Ga Gallium 69.72	32 Ge Germanium 72.61	33 As Arsenic 74.92	34 Se Selenium 78.96	35 Br Bromine 79.90	36 Kr Krypton 83.80				
37 Rb Rubidium 85.47	38 Sr Strontium 87.62	39 Y Yttrium 88.91	40 Zr Zirconium 91.22	41 Nb Niobium 92.91	42 Mo Molybdenum 95.94	43 Tc Technetium (98)	44 Ru Ruthenium 101.07	45 Rh Rhodium 102.91	46 Pd Palladium 106.42	47 Ag Silver 107.87	48 Cd Cadmium 112.41	49 In Indium 114.82	50 Sn Tin 118.71	51 Sb Antimony 121.76	52 Te Tellurium 127.60	53 I Iodine 126.90	54 Xe Xenon 131.29				
55 Cs Cesium 132.91	56 Ba Barium 137.33	57 La Lanthanum 138.91	72 Hf Hafnium 178.49	73 Ta Tantalum 180.95	74 W Tungsten 183.84	75 Re Rhenium 186.21	76 Os Osmium 190.23	77 Ir Iridium 192.22	78 Pt Platinum 195.08	79 Au Gold 196.97	80 Hg Mercury 200.59	81 Tl Thallium 204.38	82 Pb Lead 207.2	83 Bi Bismuth 208.98	84 Po Polonium (209)	85 At Astatine (210)	86 Rn Radon (222)				
87 Fr Francium (223)	88 Ra Radium (226)	89 Ac Actinium (227)	104 Rf Rutherfordium 178.49	105 Db Dubnium (262)	106 Sg Seaborgium (266)	107 Bh Bohrium (264)	108 Hs Hassium (269)	109 Mt Meitnerium (268)	110 Ds Darmstadtium (281)	111 Rg Roentgenium (272)	112 Cn Copernicium (285)										
			58 Ce Cerium 140.12	59 Pr Praseodymium 140.91	60 Nd Neodymium 144.24	61 Pm Promethium (145)	62 Sm Samarium 150.36	63 Eu Europium 151.96	64 Gd Gadolinium 157.25	65 Tb Terbium 158.93	66 Dy Dysprosium 162.50	67 Ho Holmium 164.93	68 Er Erbium 167.26	69 Tm Thulium 168.93	70 Yb Ytterbium 173.04	71 Lu Lutetium 174.97					
			90 Th Thorium 232.04	91 Pa Protactinium 231.04	92 U Uranium 238.03	93 Np Neptunium (237)	94 Pu Plutonium (244)	95 Am Americium (243)	96 Cm Curium (247)	97 Bk Berkelium (247)	98 Cf Californium (251)	99 Es Einsteinium (252)	100 Fm Fermium (257)	101 Md Mendelevium 168.93	102 No Nobelium (259)	103 Lr Lawrencium (262)					

Note: It is often confusing for students to see the terms “atom” and “element” used interchangeably as if they are the same thing. Explain to students that an atom is the smallest particle or “building block” of a substance. An element is a substance made up of all the same type of atom. For instance, a piece of pure carbon is made up of only carbon atoms. This piece of pure carbon is a sample of the element carbon. The people who developed the periodic table could have called it the Periodic Table of the Atoms but they did not have a firm understanding of atoms at that time. Since they were working with actual samples of elements such as copper, mercury, sulfur, etc., they called it the periodic table of the elements.

Optional

Play one or both of the following songs.

- *The Elements* by Tom Lehrer with animation by Mike Stanfill
www.privatehand.com/flash/elements.html
- *Meet the Elements* by They Might be Giants
www.youtube.com/watch?v=d0zION8xjbM

2. Explain the meaning of the numbers and letters in the boxes in the periodic table.

Tell students that the class will focus on the first 20 elements over 2 days. On the first day, they will look at the number of protons, electrons, and neutrons in the atoms of each element. On the second day, they will look at the arrangement of electrons in the atoms.

Give each student a copy of the periodic table of the elements, the periodic table of elements 1–20, and the activity sheet.

Students will use the periodic table of elements 1–20, along with the activity sheet, in the lesson they will do today.

Project the image *Periodic Table of the First 20 Elements*.

www.middleschoolchemistry.com/multimedia/chapter4/lesson2#first_twenty

PERIODIC TABLE ELEMENTS 1-20							
HYDROGEN 1 H 1.01							HELIUM 2 He 4.00
LITHIUM 3 Li 6.94	BERYLLIUM 4 Be 9.01	BORON 5 B 10.81	CARBON 6 C 12.01	NITROGEN 7 N 14.01	OXYGEN 8 O 16.00	FLUORINE 9 F 19.00	NEON 10 Ne 20.18
SODIUM 11 Na 22.99	MAGNESIUM 12 Mg 24.31	ALUMINUM 13 Al 26.98	SILICON 14 Si 28.09	PHOSPHORUS 15 P 30.97	SULFUR 16 S 32.07	CHLORINE 17 Cl 35.45	ARGON 18 Ar 39.95
POTASSIUM 19 K 39.10	CALCIUM 20 Ca 40.08						

Project the image *Element explanation.*

www.middleschoolchemistry.com/multimedia/chapter4/lesson2#element_explanation

Explain what the numbers and letters in each box on the periodic table represent.

<p>Atomic number ———</p> <p>The number of protons in the nucleus of the atom.</p>	<p>CARBON</p> <p>6</p> <p>C</p> <p>12.01</p>	<p>Element name</p> <p>Usually from a Greek or Latin word for the element or a substance containing the element.</p>
<p>Atomic mass ———</p> <p>The average mass of the atoms in an element.</p>		<p>Symbol</p> <p>Short-hand abbreviation for the element name.</p>

Explain atomic mass.

The atomic mass of an element is based on the mass of the protons, neutrons, and electrons of the atoms of that element. The mass of the proton and neutron are about the same, but the mass of the electron is much smaller (about 1/2000 the mass of the proton or neutron). The majority of the atomic mass is contributed by the protons and neutrons.

For any element in the periodic table, the number of electrons in an atom of that element always equals the number of protons in the nucleus. But this is not true for neutrons. Atoms of the same element can have different numbers of neutrons than protons. Atoms of the same element with different numbers of neutrons are called *isotopes* of that element. The atomic mass in the periodic table is an average of the atomic mass of the isotopes of an element. For the atoms of the first 20 elements, the number of neutrons is either equal to or slightly greater than the number of protons.

For example, the vast majority of carbon atoms have 6 protons and 6 neutrons, but a small percentage have 6 protons and 7 neutrons, and an even smaller percentage have 6 protons and 8 neutrons. Since the majority of carbon atoms have a mass very close to 12, and only a small percentage are greater than 12, the average atomic mass is slightly greater than 12.

3. Describe the activity students will do to learn about the first 20 elements of the periodic table.

Show students that you have 100 cards (5 for each of the first 20 elements). Explain that each card contains information about one of the first 20 atoms of the periodic table. The students' job is to read the card carefully, figure out which atom the card is describing, and put the card at the spot in the room for that atom.

Review the information about protons, electrons, and neutrons students need to know in order to match the cards with the correct element:

Proton

- Positively charged particle in the nucleus of the atom.
- The number of protons in an atom's nucleus is the atomic number.

Electron

- Negatively charged particle surrounding the nucleus of the atom.
- The number of electrons surrounding the nucleus of an atom is equal to the number of protons in the atom's nucleus.

Neutron

- Particle in the nucleus that has almost the same mass as a proton but has no charge.
- For the atoms of the first 20 elements, the number of neutrons is either equal to or slightly greater than the number of protons.

To match the number of neutrons listed on your card to the correct element, look for an element on the periodic table so that if you add the number of neutrons on your card to the protons of the element, you will get close to the atomic mass for that element.

For example, you may have a card that says that the atom you are looking for has 5 neutrons. You would look at the periodic table to find an atom that you could add 5 to its number of protons that would give you a sum close to the atomic mass given for that element. The answer is beryllium (Be), which has 4 protons and an atomic mass of 9.01.

Note: There are a few neutron cards that have two possible correct elements instead of just one:

- 6 Neutrons—Boron or Carbon
- 10 Neutrons—Fluorine or Neon
- 12 Neutrons—Sodium or Magnesium
- 16 Neutrons—Phosphorous or Sulfur
- 20 Neutrons—Potassium or Calcium

EXPLORE

4. Have groups work together to place each card with its correct atom.

Distribute the cards to groups. If you have 10 groups, each group will get 10 cards. Be available to help students who have trouble with the neutrons and atomic mass.



5. Discuss the placement of the cards for two or three atoms.

Select two or three atoms and review whether the cards were placed correctly. This review will help reinforce the concepts about the structure of atoms and help students determine the number of protons, electrons, and neutrons in each type of atom.

Have students begin filling out the activity sheet with the following information:

- Number of protons
- Number of electrons
- Number of neutrons (usually)

PERIODIC TABLE ELEMENTS 1-20

Write the number of protons, electrons, and neutrons in each element.

HYDROGEN 1 # of Protons: 1 # of Electrons: 1 # of Neutrons: 0 1.01							HELIUM 2 # of Protons: 2 # of Electrons: 2 # of Neutrons: 2 4.00
LITHIUM 3 # of Protons: 3 # of Electrons: 3 # of Neutrons: 4 6.94	BERYLLIUM 4 # of Protons: 4 # of Electrons: 4 # of Neutrons: 5 9.01	BORON 5 # of Protons: 5 # of Electrons: 5 # of Neutrons: 6 10.81	CARBON 6 # of Protons: 6 # of Electrons: 6 # of Neutrons: 6 12.01	NITROGEN 7 # of Protons: 7 # of Electrons: 7 # of Neutrons: 7 14.01	OXYGEN 8 # of Protons: 8 # of Electrons: 8 # of Neutrons: 8 16.00	FLUORINE 9 # of Protons: 9 # of Electrons: 9 # of Neutrons: 10 19.00	NEON 10 # of Protons: 10 # of Electrons: 10 # of Neutrons: 10 20.18
SODIUM 11 # of Protons: 11 # of Electrons: 11 # of Neutrons: 12 22.99	MAGNESIUM 12 # of Protons: 12 # of Electrons: 12 # of Neutrons: 12 24.31	ALUMINUM 13 # of Protons: 13 # of Electrons: 13 # of Neutrons: 14 26.98	SILICON 14 # of Protons: 14 # of Electrons: 14 # of Neutrons: 14 28.09	PHOSPHORUS 15 # of Protons: 15 # of Electrons: 15 # of Neutrons: 16 30.97	SULFUR 16 # of Protons: 16 # of Electrons: 16 # of Neutrons: 16 32.07	CHLORINE 17 # of Protons: 17 # of Electrons: 17 # of Neutrons: 18 35.45	ARGON 18 # of Protons: 18 # of Electrons: 18 # of Neutrons: 22 39.95
POTASSIUM 19 # of Protons: 19 # of Electrons: 19 # of Neutrons: 20 39.10	CALCIUM 20 # of Protons: 20 # of Electrons: 20 # of Neutrons: 20 40.08	<p><i>Note: The number of neutrons may be different in the atoms of the same element. The atoms of an element with different numbers of neutrons are called isotopes of that element. The number of neutrons shown in the chart represents the most common isotope for that element.</i></p>					

EXTEND

6. Introduce students to their element project and an online resource that they can use.

Assign each student to an element. Include the first 20 elements and any other elements that you find interesting so that each student can research and present their own.

Each student should find and present some basic information about their element to the class. The presentation can be in the form of a poster, pamphlet, PowerPoint presentation or other form. The presentations should be short and can include: atom name, atomic number, derivation of name, when and where discovered, natural sources of the element, major uses, and any other information you find important.

Some Internet sources for this information can be overwhelming. They can also contain advertising that you may not want students exploring. For basic information about the periodic table, including some images and video, *The Journal of Chemical Education's Periodic Table Live* is an excellent resource.

www.chemeddl.org/collections/ptl/

If there is time available, have students work on this atom project during the week.