Chapter 6, Lesson 9 Activity Sheet Answers

1. The solution became close to neutral when the color got close to the green color of the control universal indicator solution.
2. As more drops of sodium carbonate are added to the citric acid solution, the solution becomes less acidic.
3. When a base is added to an acidic solution, protons from the H_3O^+ in the solution are transferred to the base.
4. When a solution is neutral, the concentration of H_3O^+ and OH^- are equal.
5. It takes more sodium carbonate solution to neutralize a more concentrated citric acid solution.
6. In a more concentrated acid solution, there are more H_3O^+ ions in the solution. It takes more molecules of base to accept protons from these ions to make the concentration of H_3O^+ and OH^- the same.
7. Sodium carbonate solution B is more concentrated than sodium carbonate solution A. You know this because it takes more drops of citric acid solution to neutralize solution B than it takes to neutralize solution A.
8. There is probably some kind of base in the antacid medicine. The base could help neutralize some of the stomach acid.